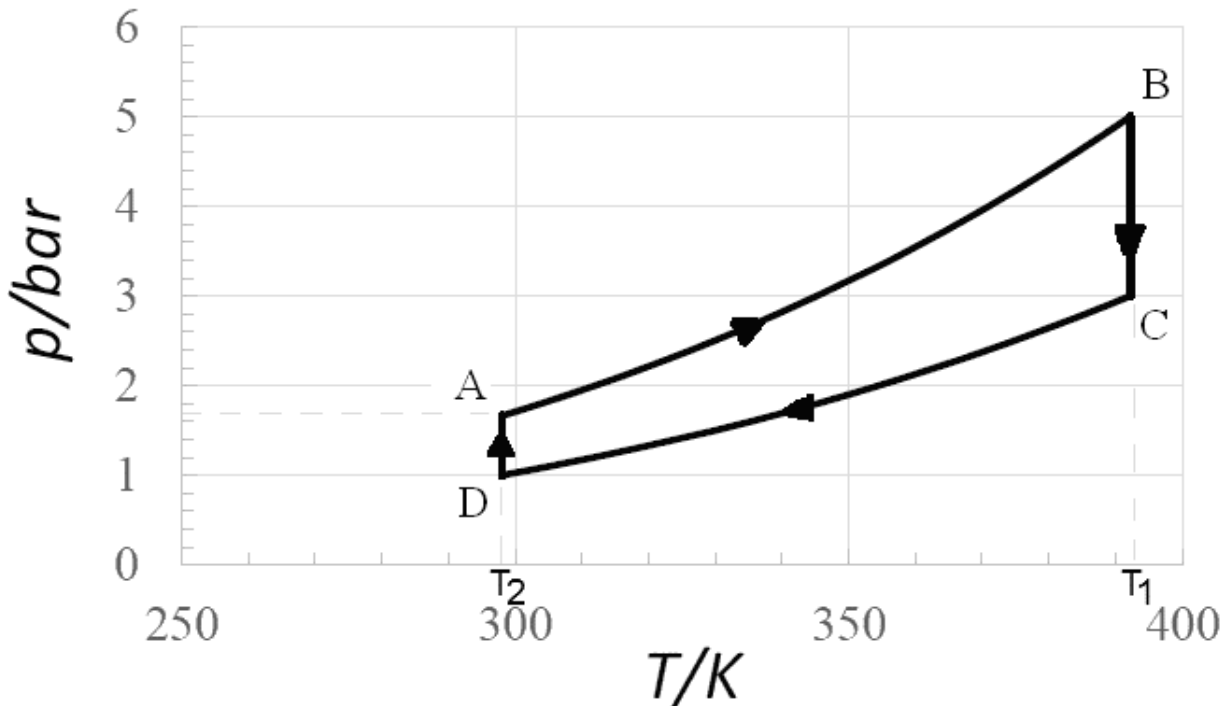


Problema 3 Carnot cycle


These data are read from the graphic:

Point A) $P_A = 1.667 \text{ bar}$; $T_2 = 298 \text{ K}$

Point B) $P_B = 5.0 \text{ bar}$; $T_1 = 392 \text{ K}$

Point D) $P_D = 1.0 \text{ bar}$; $T_2 = 298 \text{ K}$

Point C) $P_C = 3.0 \text{ bar}$; $T_1 = 392 \text{ K}$

3.1 In a thermodynamic cycle $\Delta U = 0 \Rightarrow Q = W$

in the adiabatic compression AB: $Q_{AB} = 0$

in the isothermal expansion BC: $Q_{BC} = W_{BC} = n R T_1 \ln(V_C/V_B) = n R T_1 \ln(P_B/P_C) =$
 $= n \text{ mol } 8,314 \text{ J/(mol K)} 392 \text{ K} \ln(5/3) = 1665 n \text{ J}$

in the adiabatic expansion CD: $Q_{CD} = 0$

in the isothermal compression DA: $Q_{DA} = W_{DA} = n R T_2 \ln(V_A/V_D) = n R T_2 \ln(P_D/P_A) =$
 $= n \text{ mol } 8,314 \text{ J/(mol K)} 298 \text{ K} \ln(1/1,667) = -1266 n \text{ J}$

$W = Q_{AB} + Q_{BC} + Q_{CD} + Q_{DA} = 399 n \text{ J}$

The magnitude of useful work is not definable because the quantity of substance (n) is missing.

3.2 Generally the efficiency η of a thermodynamic cycle is defined as:

$\eta = (\text{useful work})/(\text{given heat}) = W/Q_{BC} = (399 n \text{ J})/(1665 n \text{ J}) = 0,24$

even, for Carnot cycle only, the efficiency is:

$\eta = 1 - (T_2/T_1) = 1 - (298 \text{ K}/392 \text{ K}) = 0,24$

3.3 In the adiabatic expansion CD: $P_C V_C^\gamma = P_D V_D^\gamma \Rightarrow P_C (T_1/P_C)^\gamma = P_D (T_2/P_D)^\gamma \Rightarrow$
 $P_C^{1-\gamma} T_1^\gamma = P_D^{1-\gamma} T_2^\gamma \Rightarrow (P_C/P_D)^{1-\gamma} = (T_2/T_1)^\gamma \Rightarrow (1-\gamma) \ln(P_C/P_D) = \gamma \ln(T_2/T_1) \Rightarrow$
 $1/\gamma - 1 = \ln(T_2/T_1)/\ln(P_C/P_D) \Rightarrow 1/\gamma = \ln(T_2/T_1)/\ln(P_C/P_D) + 1 \Rightarrow$
 $1/\gamma = \ln(298/392)/\ln(3/1) + 1 = 0,75 \Rightarrow \gamma = 1/0,75 = 1/(3/4) = 1,33$
 $\gamma = 1 + R/C_V \Rightarrow \gamma - 1 = R/C_V \Rightarrow C_V = R/(\gamma - 1) = R/(1,33 - 1) \Rightarrow C_V = 3 R$

3.4 Remembering that, for an ideal gas, $C_{V \text{ monoatomic}} = 3/2 R$, $C_{V \text{ diatomic}} = 3/2 R + 2/2 R = 5/2 R$,
 $C_{V \text{ triatomic}} = 3/2 R + 3/2 R = 3 R$, we can affirm that this molecule with $C_V = 3 R$ is a triatomic non-linear gas molecule.

Solution proposed by Roberto Tinelli – Student at ITIS "L. dell'Erba" – Castellana Grotte – Italy